

OXFORD CAMBRIDGE AND RSA EXAMINATIONS
GCSE (9–1)
J248/02

CHEMISTRY A (GATEWAY SCIENCE)
Paper 2, C4–C6 and C7
(Foundation Tier)

WEDNESDAY 13 JUNE 2018: Morning
TIME ALLOWED: 1 hour 45 minutes
plus your additional time allowance
MODIFIED ENLARGED 24pt

First name						Last name					
Centre number						Candidate number					

YOU MUST HAVE:

a ruler (cm/mm)

the Data Sheet (for GCSE Chemistry A)

YOU MAY USE:

a scientific or graphical calculator

an HB pencil

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS

The data sheet will be found with this document.

Use black ink. You may use an HB pencil for graphs and diagrams.

Complete the boxes on the front page with your name, centre number and candidate number.

Answer ALL the questions.

Write your answer to each question in the space provided. If additional space is required, use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.

INFORMATION

The total mark for this paper is 90.

The marks for each question are shown in brackets [].

Quality of extended responses will be assessed in questions marked with an asterisk (*).

SECTION A

Answer ALL the questions.

You should spend a maximum of 30 minutes on this section.

1 Lithium, sodium and potassium are Group 1 elements.

What happens when these elements are added to water? [1]

- A Some float and carbon dioxide gas and an alkaline solution are made.
- B Some float and hydrogen gas and an alkaline solution are made.
- C They all float and hydrogen gas and an acidic solution are made.
- D They all float and hydrogen gas and an alkaline solution are made.

Your answer

☐

2 Damp litmus paper is used to test for chlorine gas.

Which statement describes the correct result of the test for chlorine gas? [1]

- A Damp blue litmus paper turns red then white.
- B Damp blue litmus paper turns white then red.
- C Damp red litmus paper turns blue then white.
- D Damp red litmus paper turns white then blue.

Your answer

☐

3 What is the name of the process that converts large alkane molecules into smaller alkane molecules? [1]

A Cracking

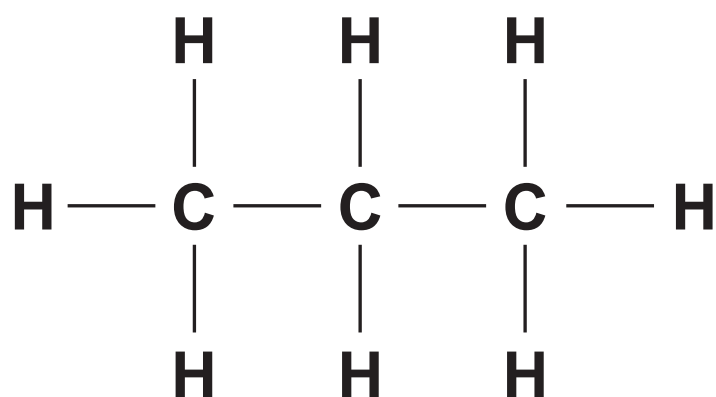
B Fractional distillation

C Hydrogenation

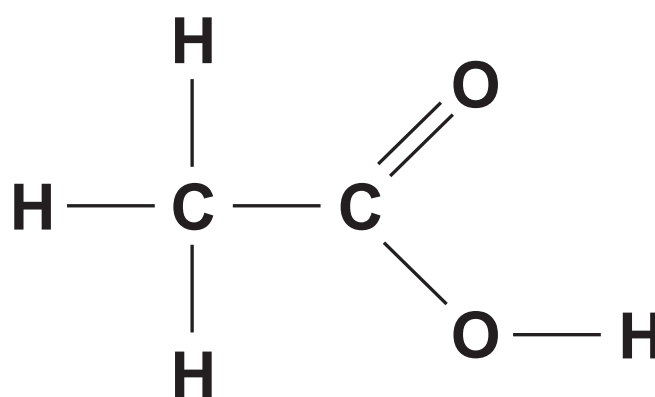
D Polymerisation

Your answer

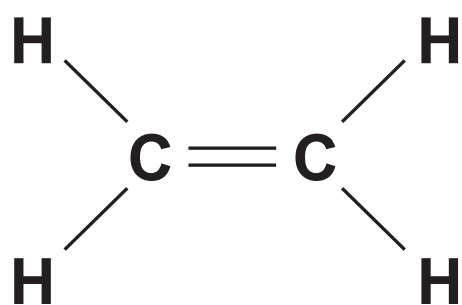
4 Which displayed formula shows an alkene? [1]



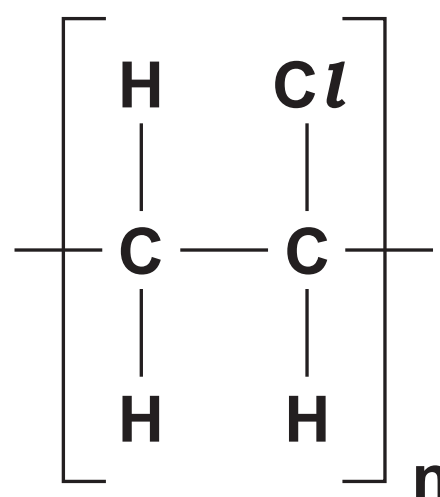
A



B



C



D

Your answer

5 The list shows part of the reactivity series of metals including carbon.

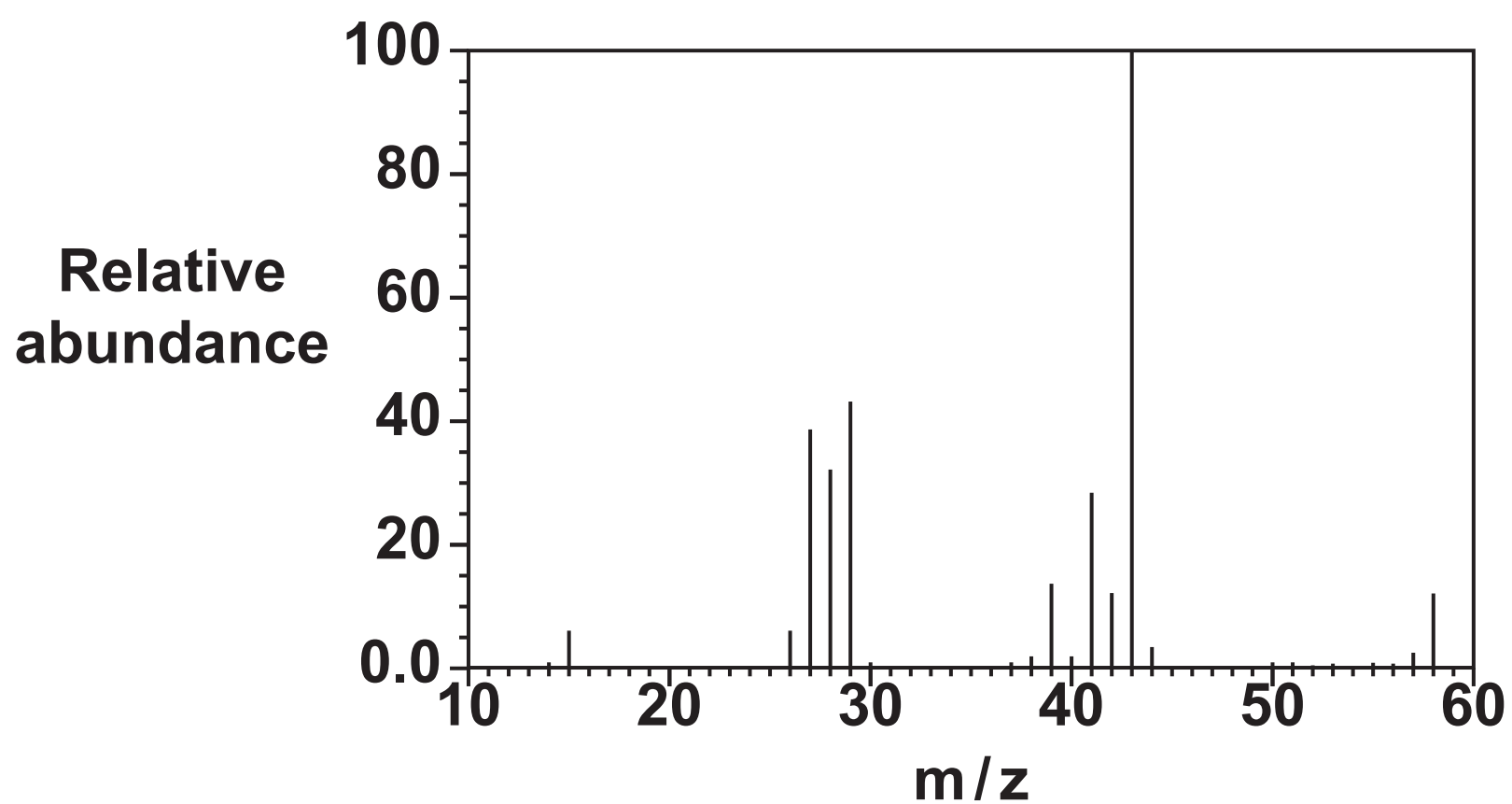
Sodium
Lithium
Calcium
Magnesium
Aluminium
Carbon
Zinc
Iron
Tin
Lead

Which row of the table correctly describes how the metals are extracted from their ores? [1]

	Metals extracted by electrolysis	Metals extracted by heating with carbon
A	Sodium, Magnesium, Zinc	Aluminium, Iron, Tin
B	Aluminium, Zinc, Iron	Lead, Tin
C	Calcium, Magnesium, Aluminium	Sodium, Iron, Tin
D	Sodium, Calcium, Magnesium	Lead, Tin, Zinc

Your answer

6 Look at the mass spectrum of a carbon compound.



Which carbon compound is the mass spectrum from? [1]



Your answer

7 Which statement about catalysts is correct? [1]

- A A catalyst decreases the activation energy of a reaction.**
- B A catalyst increases the activation energy of a reaction.**
- C A catalyst increases the time for a reaction to go to completion.**
- D A catalyst slows down a reaction.**

Your answer

8 Hydrogen gas can be made by reacting methane and steam (H₂O).



6 g of hydrogen gas can be made from 18 g of steam, H₂O.

How much hydrogen gas can be made from 3.6 g of steam, H₂O? [1]

- A 0.4 g**
- B 0.6 g**
- C 1.2 g**
- D 6.8 g**

Your answer

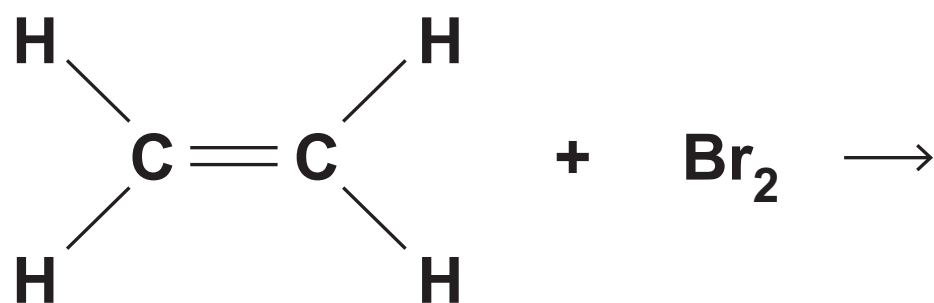
9 Which statement is true for a reversible reaction when it is at dynamic equilibrium? [1]

- A The concentration of the products is increasing.
- B The rate of the backward reaction is greater than the rate of the forward reaction.
- C The rate of the forward reaction is equal to the rate of the backward reaction.
- D The rate of the forward reaction is greater than the rate of the backward reaction.

Your answer

☐

10 What is the formula of the product in this equation? [1]

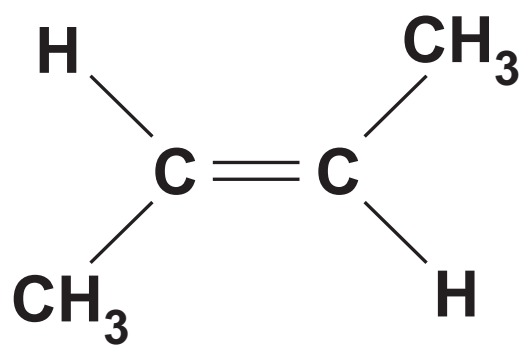


- A $\text{C}_2\text{H}_2\text{Br}$
- B C_3Br_4
- C $\text{C}_2\text{H}_3\text{Br}$
- D $\text{C}_3\text{H}_4\text{Br}_2$

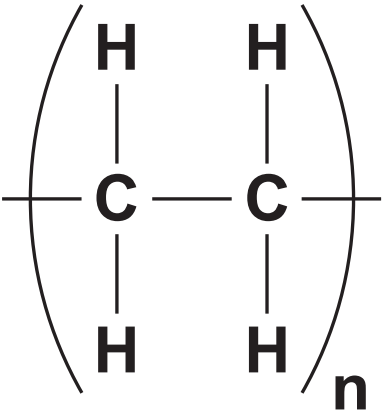
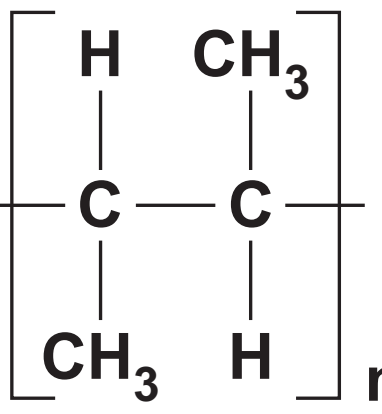
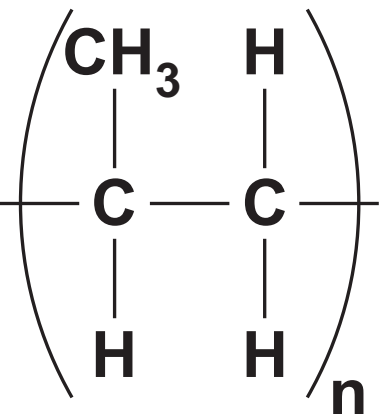
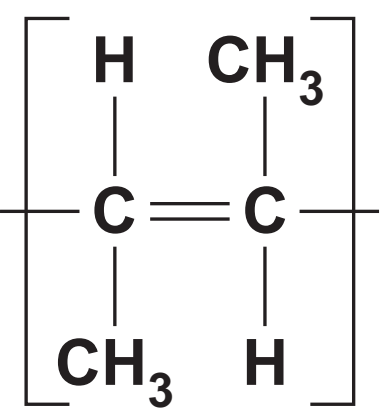
Your answer

☐

11 Look at the displayed formula of the monomer butene.



What is the formula of the POLYMER formed from butene? [1]

A	B	C	D
			

Structure A: A repeating unit in brackets with a subscript 'n'. It shows two carbon atoms connected by a single bond. Each carbon is bonded to two hydrogen atoms (one above, one below). The brackets have horizontal bonds extending from the carbons.

Structure B: A repeating unit in brackets with a subscript 'n'. It shows two carbon atoms connected by a single bond. The left carbon is bonded to a hydrogen atom (H) above and a methyl group (CH₃) below. The right carbon is bonded to a methyl group (CH₃) above and a hydrogen atom (H) below. The brackets have horizontal bonds extending from the carbons.

Structure C: A repeating unit in brackets with a subscript 'n'. It shows two carbon atoms connected by a single bond. The left carbon is bonded to a methyl group (CH₃) above and a hydrogen atom (H) below. The right carbon is bonded to a hydrogen atom (H) above and a hydrogen atom (H) below. The brackets have horizontal bonds extending from the carbons.

Structure D: A repeating unit in brackets with a subscript 'n'. It shows two carbon atoms connected by a double bond. The left carbon is bonded to a hydrogen atom (H) above and a methyl group (CH₃) below. The right carbon is bonded to a methyl group (CH₃) above and a hydrogen atom (H) below. The brackets have horizontal bonds extending from the carbons.

Your answer

12 DNA molecules are polymers.

What are the monomers that make up DNA called? [1]

A Carbohydrates

B Nucleotides

C Phosphates

D Proteins

Your answer

13 What is the major source of carbon monoxide in the Earth's atmosphere? [1]

A Incomplete combustion of fossil fuels and wood.

B Production in a nuclear power station.

C The combustion of impurities in coal.

D The combustion of impurities in natural gas.

Your answer

14 Look at the following sentences.

They describe one possible theory for how the Earth's atmosphere evolved.

The sentences are not in the correct order.

1	Carbon cycle now keeps the composition of the atmosphere almost constant
2	Initial atmosphere of ammonia and carbon dioxide
3	Increase in oxygen and nitrogen levels
4	Photosynthetic organisms began to make oxygen
5	Degassing from the Earth's crust and formation of water

What is the correct order for the sentences? [1]

A 2, 4, 3, 5, 1

B 2, 5, 4, 3, 1

C 5, 2, 3, 4, 1

D 5, 2, 4, 3, 1

Your answer

15 Look at the information about four different polymers.

Polymer	Cost (£ per kg)	Tensile strength (MPa)	Melting point (°C)	Maximum useable temperature (°C)
A	0.74	15	120	85
B	1.20	78	254	70
C	0.92	35	176	160
D	1.42	42	156	160

Which polymer would be best for making a plastic cup to hold hot drinks? [1]

Your answer

SECTION B

Answer ALL the questions.

16 The Group 7 elements are called the halogens.

The table shows information about some of the halogens.

Name	Atomic number	Boiling point (°C)	State at room temperature	Molecular formula
Fluorine	9	−188	Gas	F ₂
Chlorine	17	−34	Gas	Cl ₂
Bromine	35	59	Liquid	Br ₂
Iodine	53	184	Solid	I ₂

(a) Which is the MOST REACTIVE halogen in the table?

_____ [1]

(b) Astatine is also a halogen. It has the atomic symbol At and an atomic number of 85.

Look at the table.

(i) Predict the STATE of astatine at room temperature.

_____ [1]

(ii) Predict the BOILING POINT of astatine.

_____ [1]

(c) Sodium, Na, reacts with chlorine. A white solid is made.

(i) What is the NAME of this white solid?

_____ **[1]**

(ii) Write down the BALANCED SYMBOL equation for this reaction.

_____ **[2]**

- 17 A student added 2.4 g of magnesium to hydrochloric acid. She observed that no magnesium was left when the reaction was complete.**

The student transferred the solution to an evaporating basin. She heated the solution using a Bunsen burner and evaporated all the water.

- (a) Explain how you can tell from the student's observation that the hydrochloric acid was in excess.**

[1]

- (b) Look at the equation for the reaction.**



The student knows the reaction is complete when there is no magnesium left.

Use the equation to explain one OTHER way the student could tell that the reaction was complete.

[1]

(c) The student predicts she should make 9.5 g of magnesium chloride, MgCl_2 .

She actually makes 7.9 g.

Calculate the PERCENTAGE YIELD.

Give your answer to 3 significant figures.

Answer = _____ [3]

(d) Write down one reason, other than a mistake, why the student may have obtained a percentage yield of less than 100%.

_____ [1]

18 Antacid tablets are used to treat indigestion.

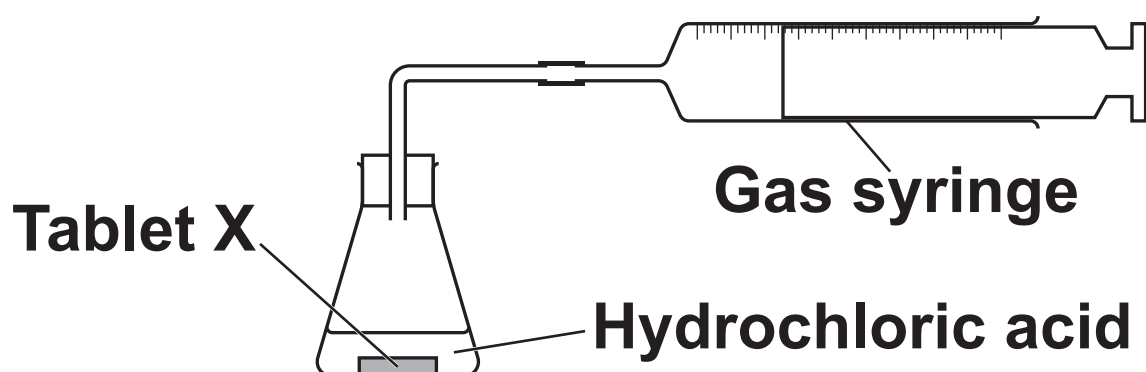
A student investigates two different antacid tablets, X and Y. Both tablets, X and Y, contain calcium carbonate, CaCO_3 .

Calcium carbonate reacts with hydrochloric acid. Calcium chloride, CaCl_2 , water and carbon dioxide are made.

(a) Write a BALANCED SYMBOL equation for this reaction.

_____ [2]

(b) The diagram shows the apparatus the student uses.



The student reacts tablet X with 100 cm^3 of hydrochloric acid. The hydrochloric acid is in excess.

He measures the volume of gas made every minute during the first five minutes.

He does a second experiment using tablet Y and a fresh 100 cm^3 sample of the same hydrochloric acid.

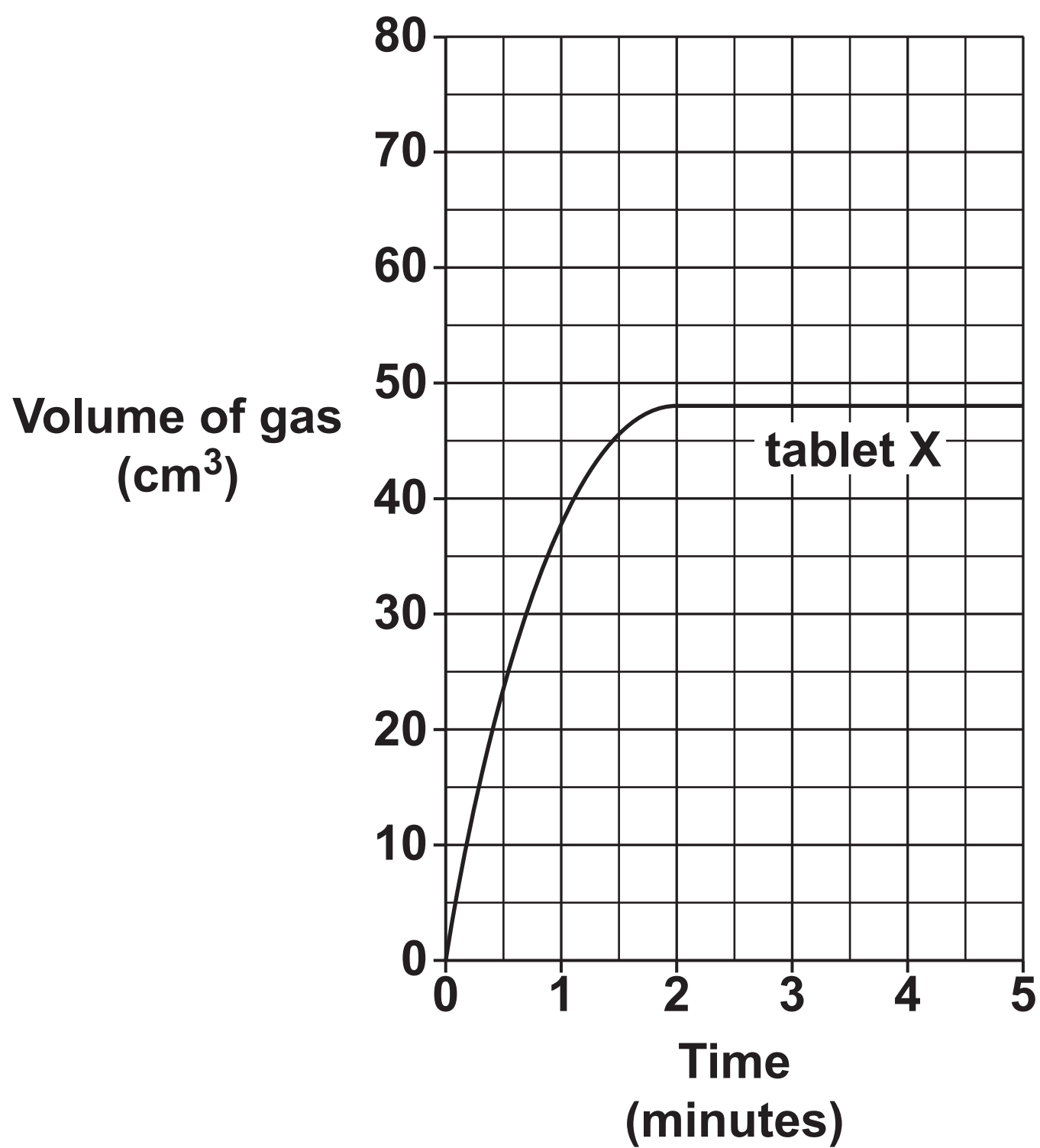
The table shows his results.

Time (minutes)	Volume of gas (cm^3)	
	Tablet X	Tablet Y
0	0	0
1	38	32
2	48	54
3	48	67
4	48	72
5	48	72

(i) The graph opposite shows the results for tablet X.

What is the volume of gas made by the end of the experiment?

Answer = _____ cm^3 [1]



(ii) Plot the results for tablet Y on the grid. Draw a line of best fit. [2]

(iii) Tablet X contains less calcium carbonate than tablet Y.

How do the results show this?

[1]

(c) The rate of reaction between calcium carbonate and hydrochloric acid can be increased by:

Using a more concentrated solution of hydrochloric acid

Increasing the temperature of the acid.

Explain how each of these methods increase the rate of the reaction.

Use ideas about collisions between particles.

[4]

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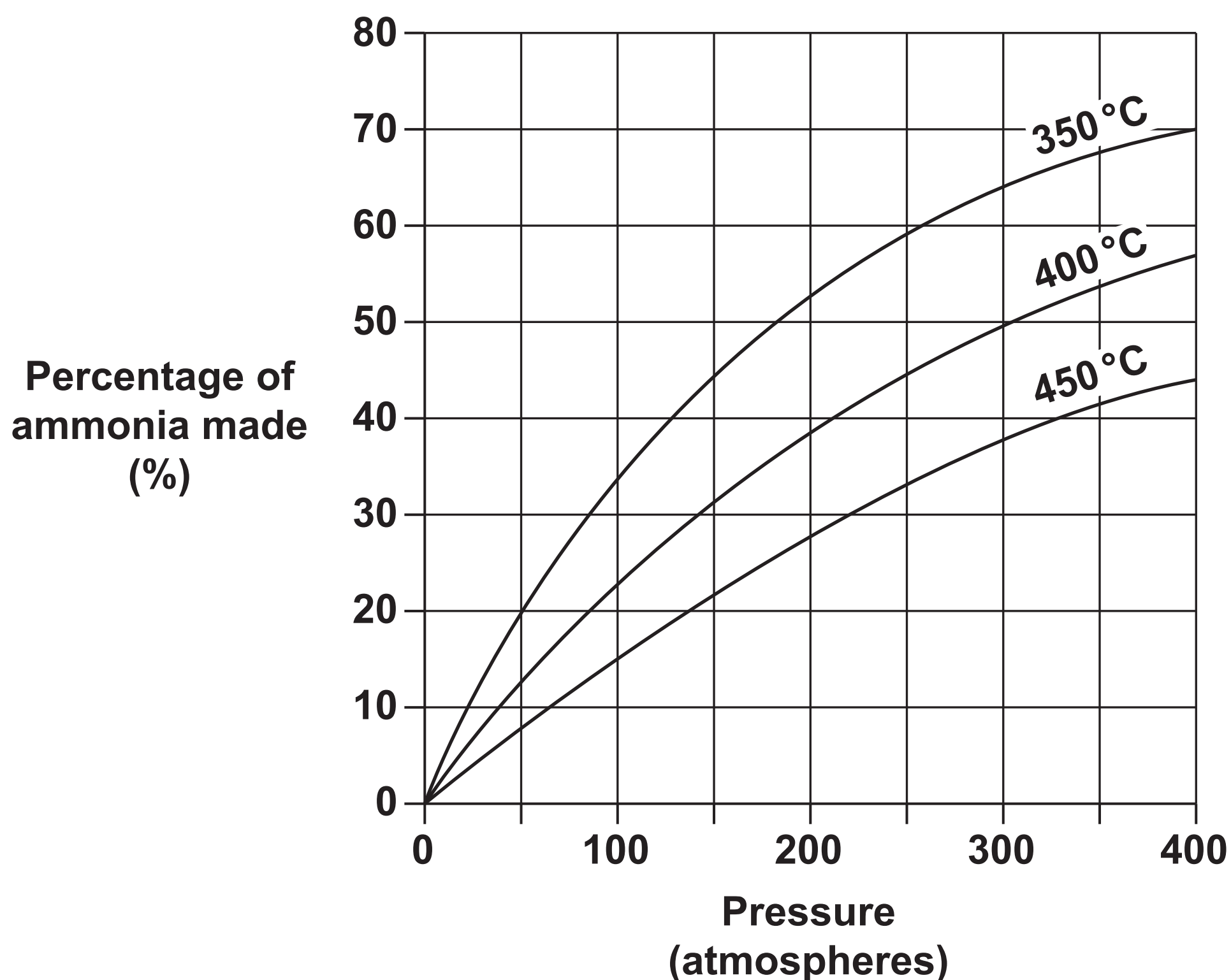
19 Ammonia is made from nitrogen and hydrogen.

- (a) The reaction between nitrogen and hydrogen is **REVERSIBLE**.

Explain what this means.

[1]

- (b) The graph shows the percentage of ammonia made at different temperatures and pressures.



- (i) Describe how the percentage of ammonia changes as the PRESSURE increases at 450 °C.

_____ [1]

- (ii) Write down a temperature and pressure which make 20% of ammonia. [1]

Temperature = _____ °C

Pressure = _____ atmospheres

20 A student wants to identify the ions contained in a solid, X.

She dissolves the solid in some water and then does some tests on the solution.

Look at the table of her results.

Test	Method	Observations
Test 1	Flame test	Red flame seen
Test 2	Add dilute sodium hydroxide solution	White precipitate forms which re-dissolves in excess sodium hydroxide solution
Test 3	Add dilute nitric acid, then silver nitrate solution	White precipitate forms
Test 4	Add dilute hydrochloric acid, then barium chloride solution	No change – mixture stays clear and colourless

(a) Describe how the student does the flame test in TEST 1.

You may draw a LABELLED diagram to help your answer.

[3]

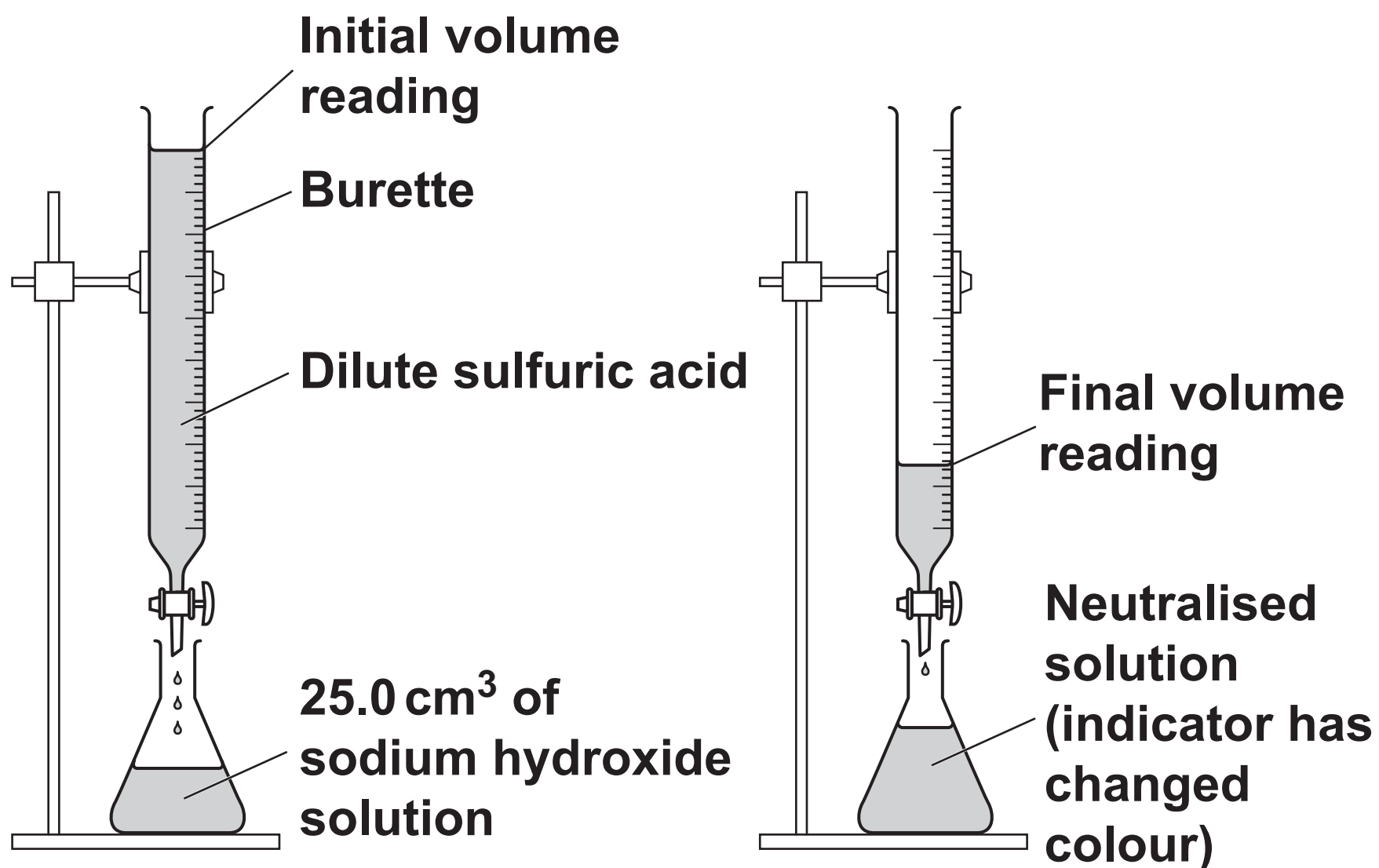
(b)*The student thinks that solid X contains only lithium ions, Li^+ , and sulfate ions, SO_4^{2-} .

Use her results to explain if she is correct.

[6]

21 A student does a titration with an acid and an alkali.

He uses dilute sulfuric acid, sodium hydroxide solution and an indicator solution.



The student's method is:

Use a measuring cylinder to pour 25.0 cm³ of sodium hydroxide solution into a conical flask

Add a few drops of an indicator to the sodium hydroxide solution

Use a burette to add dilute sulfuric acid to the sodium hydroxide solution until the indicator changes colour.

- (a) The student wants to get a more accurate value for how much acid reacts with 25.0 cm^3 of sodium hydroxide solution.

Describe and explain how the student could improve his experiment to get a more accurate value.

[4]

- (b) Another student does a titration. She also uses dilute sulfuric acid, sodium hydroxide solution and an indicator solution.

The table shows her results.

Titration number	1	2	3	4
Volume of acid (cm^3)	26.4	25.2	25.6	25.4

The student decides that the best value for the mean (average) volume of acid is 25.4 cm^3 .

Show how she calculated this value.

[2]

(c) The equation for this reaction is



	Relative formula mass, M_r
H_2SO_4	98
NaOH	40
Na_2SO_4	142
H_2O	18

Water is a waste product in this reaction.

Calculate the ATOM ECONOMY for the reaction.

Give your answer to 1 decimal place.

Answer = _____ [3]

22 Scientists are worried about the greenhouse effect.

(a) Complete the following paragraph about the greenhouse effect. [2]

Use words from the list.

- CFCs
- cool
- Earth’s
- hydrogen
- infrared
- methane
- microwave
- Sun’s
- ultraviolet
- warm

Greenhouse gases such as carbon dioxide and _____ absorb _____ radiation radiated by the _____ surface, then emit it in all directions. This greenhouse effect keeps the Earth and its atmosphere _____ enough for living things to exist.

(b) Write about TWO possible ways that greenhouse gas emissions can be reduced.

[2]

23 This question is about metals and alloys.

(a) The table gives information about some alloys.

Alloy	Main metal or metals	Use
Brass		Musical instruments and coins
Bronze		Statues
Duralumin		Aircraft parts
Solder	Lead and tin	Joining metals
Steel	Iron	Bridges, cars

Complete the table.

Choose your answers from the list. [3]

Aluminium and copper

Aluminium and iron

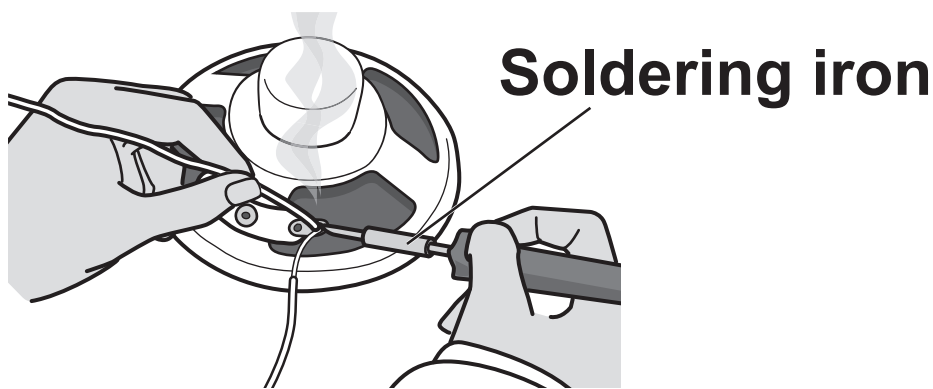
Copper and tin

Copper and zinc

Copper and lead

Lead and zinc

(b) Solder can be used to join metals together. A hot soldering iron is used to melt the solder.



The table gives some information about solder, copper and tin.

Metal	Melting point (°C)	Density (g/cm ³)	Relative hardness
Copper	1085	8.96	Soft
Tin	232	7.31	Soft
Solder	130	10.3	Quite hard

Solder is better than copper or tin for joining metals together.

Suggest why. Use the information in the table.

[2]

(c) Steel is an alloy containing iron.

Complete the WORD EQUATION for the corrosion of iron.
[2]

Iron + _____ + _____
→ _____

(d) (i) Iron can be plated with a layer of ZINC to prevent it corroding.

This is called GALVANISING.

Explain how galvanising prevents iron from corroding.

_____ **[2]**

(ii) Iron can also be plated with a layer of TIN to prevent it corroding.

Describe a DISADVANTAGE of tin plating for preventing corrosion.

_____ **[1]**

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24 This question is about life-cycle assessment.

(a) A car company is developing three new cars:

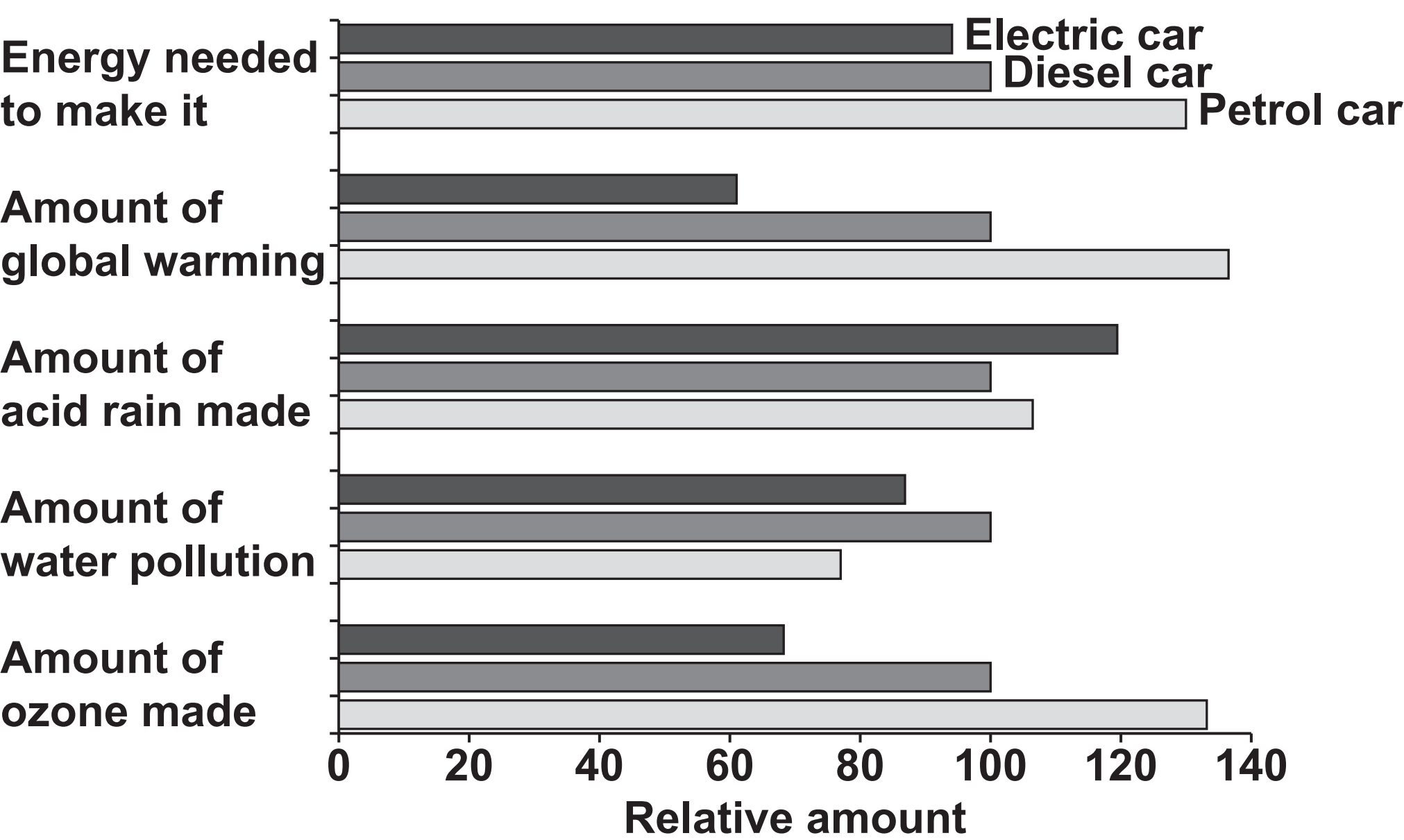
A petrol car

A diesel car

An electric car.

They do a life-cycle assessment of each car.

Look at the information about the life-cycle assessment of each car.



The company decides to manufacture and sell the electric car.

Explain why they make this choice.

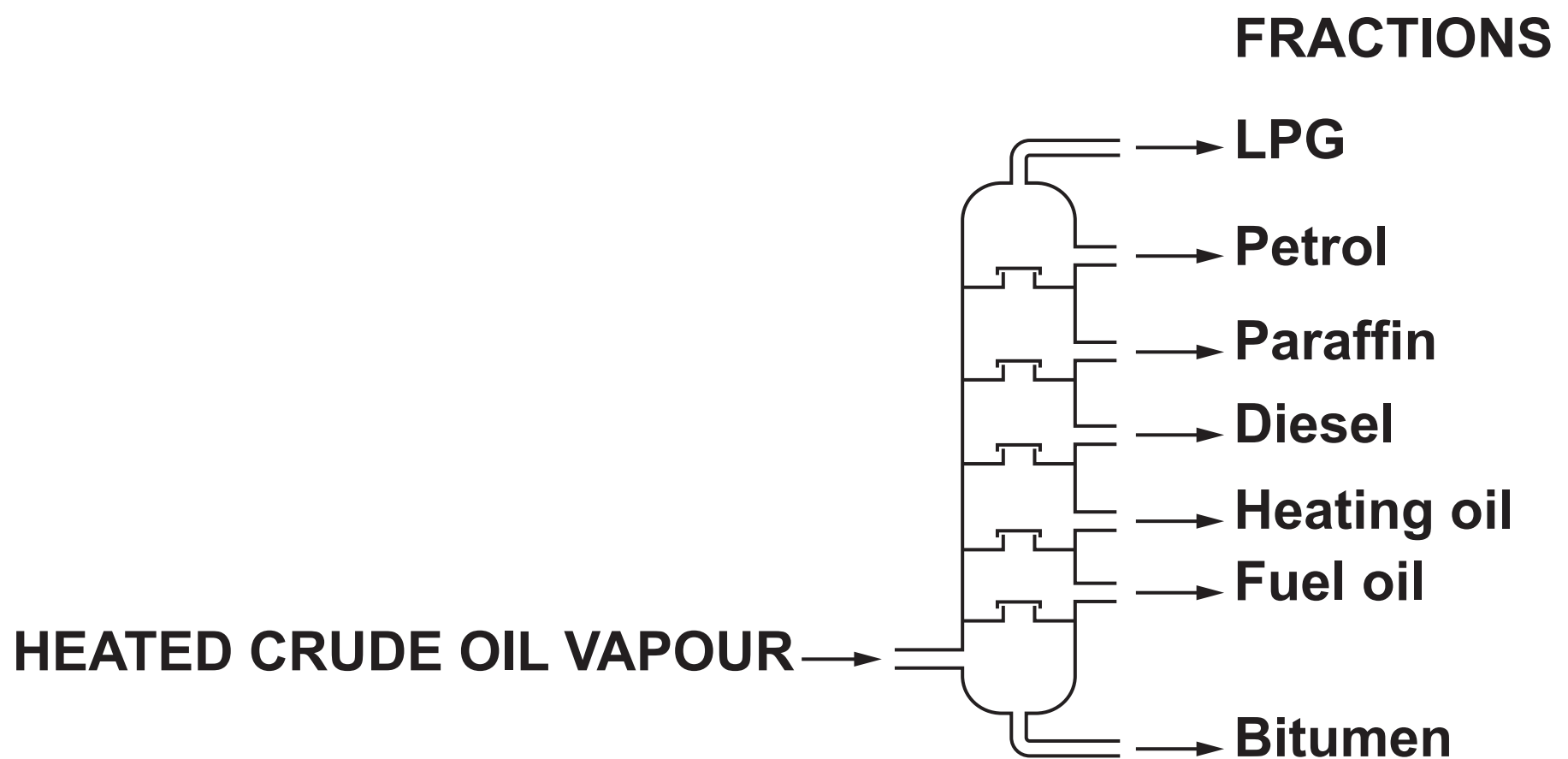
Use the information from the life-cycle assessment to help you.

[3]

(b) The fuels for the petrol and diesel cars are made from crude oil.

Crude oil is separated into different parts by **FRACTIONAL DISTILLATION**.

The diagram shows a fractionating column.



Explain why crude oil VAPOUR can be separated by fractional distillation.

[3]

(c) The table shows the boiling points of molecules present in different crude oil fractions.

Molecule	Boiling point (°C)
A	-2
B	125
C	216
D	502

Which molecule, A, B, C or D, is in the LPG FRACTION?

Explain your decision.

[2]

(d) Car manufacturers are developing cars that are powered by hydrogen/oxygen fuel cells.

The table shows some information about a 200 km journey using an electric car and a car using a fuel cell.

Feature	Electric	Fuel cell
Refuelling time (minutes)	360	4
Cost of refuelling (£)	3.20	4.20
CO ₂ emitted (kg)	48	36
Mass of car (kg)	1550	1200

Evaluate the **ADVANTAGES** and **DISADVANTAGES** of using a car powered by a fuel cell, rather than an electric car for the 200 km journey.

[3]

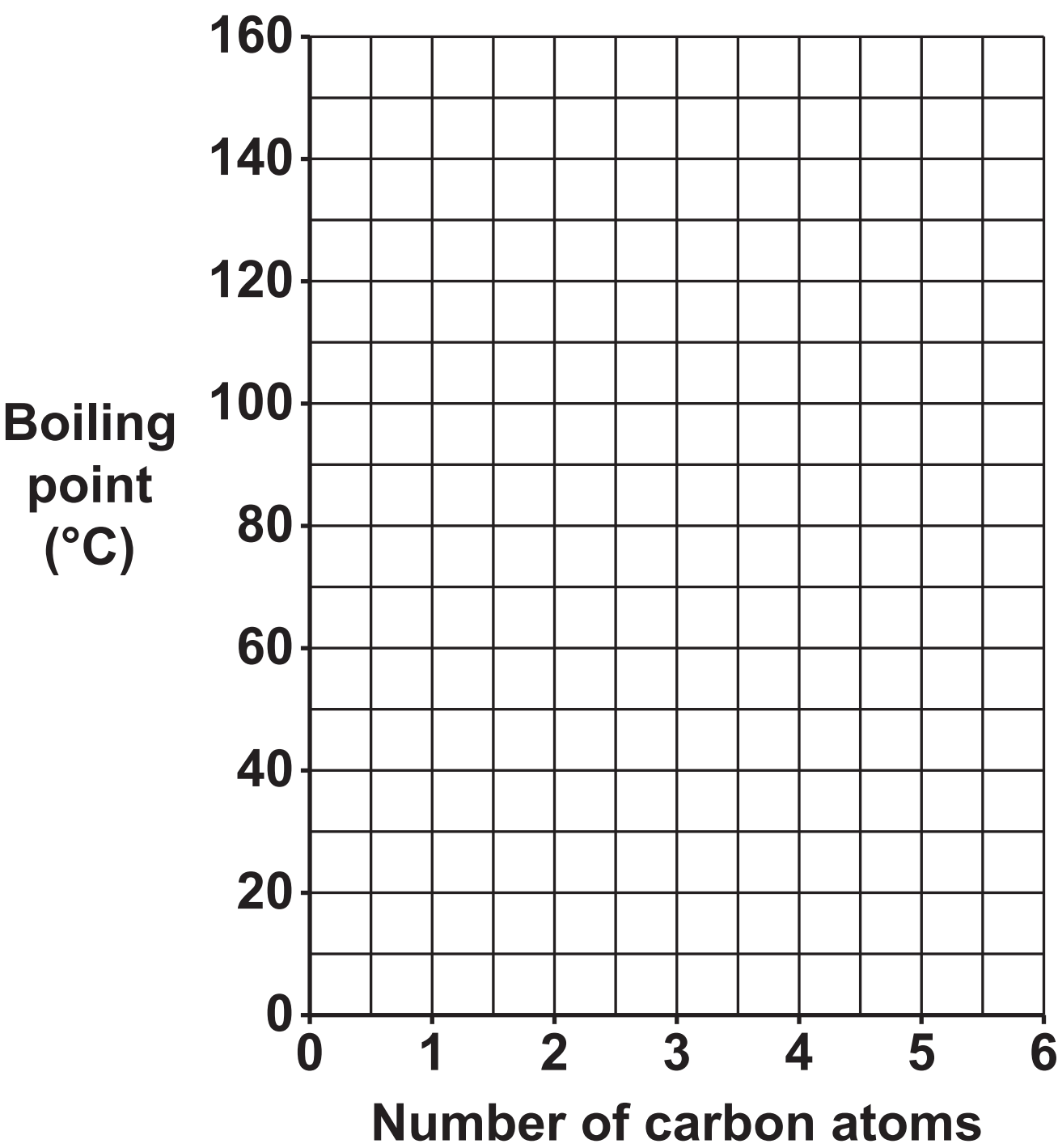
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25 A student is using the internet to find out about alcohols. The student finds the following information.

Name	Number of carbon atoms	Boiling point (°C)
Methanol	1	65
Ethanol	2	79
Propanol	3	97
Pentanol	5	138
Hexanol	6	156

(a) Plot a graph of the boiling points of the alcohols on the grid. Draw a line of best fit. [3]



- (b) (i) The student could not find a value for the boiling point of butanol, $\text{C}_4\text{H}_9\text{OH}$.

Use the graph to estimate the boiling point of butanol.

Answer = _____ °C [1]

- (ii) Draw the **DISPLAYED FORMULA** of butanol, $\text{C}_4\text{H}_9\text{OH}$.
[1]

- (c) The alcohols all react in a similar way because they all contain the same **FUNCTIONAL GROUP**.

What is the functional group in an **ALCOHOL** molecule?

_____ [1]

- (d) Ethanol, $\text{C}_2\text{H}_5\text{OH}$, can be oxidised to **ETHANOIC ACID** using potassium manganate(VII).

What is the formula of ethanoic acid?

_____ [1]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

This image shows a blank sheet of white paper with horizontal ruling lines. A single vertical line runs down the left side, creating a margin. There are 20 horizontal lines in total, evenly spaced across the page. The lines are thin and black.

